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defined, and the  
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between solvent  
and solute  
components is  
stated. Several  
examples of  
mixtures that  
are ...

pH, pOH, H<sub>3</sub>O<sup>+</sup>,  
OH<sup>-</sup>, K<sub>w</sub>, K<sub>a</sub>, K<sub>b</sub>,  
pK<sub>a</sub>, and pK<sub>b</sub>

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*pH, pOH, H<sub>3</sub>O<sup>+</sup>,  
OH<sup>-</sup>, K<sub>w</sub>, K<sub>a</sub>, K<sub>b</sub>,  
pK<sub>a</sub>, and pK<sub>b</sub>*

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, .



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(Ksp)

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*Professor Dave*

*Explains vor 1*

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*36 Sekunden*

*76.634 Aufrufe*

*We've learned*

*that some ionic*

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*solids are  
totally water  
insoluble, but  
in fact this is  
a slight oversim  
plification.  
Even such solids  
will ...*

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**Solubility can**  
**be difficult if**  
**you don't know**  
**how to properly**  
**use a solubility**  
**table! In this**  
**video I go over**  
**all of the**  
**solubility**  
**rules ...**

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*mixtures and ,  
solutions , ,  
but knowing  
which is which  
can be  
difficult, but  
after 2 minutes  
with me,  
difficult no ...*

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